

MICROCOPY RESOLUTION TEST CHART NATIONAL BUREAU OF STANDARDS-1963-A

OFFICE OF NAVAL RESEARCH/CHEMISTRY PROGRAM

Contract NO0014-79-C-0668

Task No. NR 056-721

TECHNICAL REPORT NO. 4

Phase Transformation Study of Cu₂S.

ЬУ

John A. Gardner, D. K. Gaskill, M. Hirscher, H. Jaeger, K. S. Krane

and R. L. Rasera

Prepared for Publication

in the

Proceedings of the 1983 International Conference on Hyperfine Interactions, Holland 1983

Oregon State University Department of Physics Corvallis, Or.

September 1, 1983



Reproduction in whole or in part is permitted for any purpose of the United States Government

DESCRIPTION STATES A

Approved for public releases Distribution Unlimited

PRASE TRANSFORMATION STUDY OF Cu 2S

John A. Gardner, D. K. Gaskill, M. Hirscher*, H. Jaeger, K. S. Krane, and R. L. Rasera**

Department of Physics, Oregon State University, Corvallis, OR 97331 USA

The phase transformations of crystalline Cu₂S in the temperature range 50°C to 600°C have been studied through electric quadrupole interactions of 111Cd impurities observed by time-differential perturbed angular correlations. The Cd is subject to no discernable quadrupolar interaction in the cubic phase above 430°C. In the hexagonal phase between approximately 103°C and 430°C, a singlesite, weakly temperature dependent, axially symmetric quadrupolar interaction is found. At lower temperature, the structure is complex with a number of different sites. These data were taken with a microcomputer-controlled TDPAC spectrometer with logical steering that allows a standard 4-detector spectrometer to collect simultaneously 12 channels of data.

1. Introduction

The crystal structures of copper-sulfur alloys near the Cu₂S composition have been difficult to determine. There are a number of distinct tu B compounds between x=1.8 and x=2.0. [1,2] For Cu₂S, the crystal structure of two high temperature phases has been measured, but the exact structure of the complex monoclinic root temperature it cell is not known. [3,4] The phase transition from the "high chalcocite" hexagonal phase to the captex "low chalcocite" phase occurs at 103.5°C for x=2.000 but is reported to drop to 90°C for the slightly more sulfur-rich compound with x=1.990. [2] Supercooling of the higher-T phase by as much as 20°C has also been noted. [2,5]

In this experimental study, the time differential perturbed angular correlation (TDPAC) spectra of gamma rays emitted by dilute 111Cd impurities in Cu₂S are measured. [6] The Cd atoms apparently enter the lattice substitutionally for Cu. Above 430°C, the lattice is cubic, and copper nuclei are in sites which have zero electric field gradient. As expected, the angular correlation of the gamma rays emitted by Cd nuclei is unperturbed in samples above 430°C. Between approximately 103°C and 430°C, the lattice is hexagonal, and the Cd TDPAC spectrum is typical of an axially symmetric quadrucole interaction. Below 103°C, the

*Present address: MaxeChanck Enstitute fur Metallforschung, 7000 Stuttgart 80, W. Germany Marie 1985

**Permanent address: Department of Physics, University of Maryland Baltimore County, Baltimore MD

Tachaina Danaut HA	4	D. J. RECIPIENT'S CATALO	5 NUMBER
Technical Report #4	AD-A13558		
L TITLE (and Subtitle)	TITLE (and Subdite)		ERIOD COVERED
Phase Transformation Study of Cu ₂ S.		Interim	
		6. PERFORMING ORG. RE	PORT NUMBER
John A. Gardner, D. K. Gaskill, M. Hirscher, H. Jaeger, K. S. Krane and R. L. Rasera		8. CONTRACT OR GRANT	NUMBER(e)
		N00014-79-C-066	8
. PERFORMING ORGANIZATION NAME AND AC	ORESS	10. PROGRAM ELEMENT.	PROJECT, TASK
Oregon State University Corvallis, OR 97331		NR-056-721	ŀ
1. CONTROLLING OFFICE NAME AND ADDRES	18	12. REPORT DATE	 .
Office of Naval Research/Che	mistry Program	September 1. 19	83
Arlington, VA 22217		13. NUMBER OF PAGES	1.
4. MONITORING AGENCY NAME & ADDRESS(IL	different from Controlling Office)	18. SECURITY CLASS. (of	thie report)
		Unclassified	
		184 DECLASSIFICATION	DOWNGRADING
6. DISTRIBUTION STATEMENT (of this Report)			Accession For
Approve for public release;	distribution unlimit	ed	NTIS GRALI
			DTIC TAB
			Unannounced Justification
7. DISTRIBUTION STATEMENT (of the abovest	entered in Block 20, if different i	rem Report)	- Justiniani
			Ву
			Distribution,
			Availability
6. SUPPLEMENTARY HOTES			A Avail a
To be published in Proceedin Hyperfine Interactions, Holl		national Conference	e on
A. KEY WORKS (Continue on reverse side if neces	•	•	
Cu ₂ S, time-differential pert	urbed angular correl	ations, phase tran	sformations.
A ABSTRACT Common on source side of second line phase transformations of c	port and identity by block member	he temperature ran	ne soer to
500°C have been studied throug	h electric quadrupol	e interactions stu	died through
electric quadrupole interaction	ns of ^{lll} Cd impuriti	ies observed by tim	e-differen-
tial perturbed angular correla	tions. The Cd is su	bject to no discer	nable quad-
rupolar interaction in the cub	1c phase above 430°C	. In the hexagona	I phase be-
busen approximately 103°C and	interaction is found	, weakly temperatu	re dependent
axially symmetric quadrupolar structure is complex with a nu	mber of different si	tes. These data w	ere taken
- 2000	I OSCOLETE		
S/N 9192-LF-014-0001	SECURITY CL	ASSIFICATION OF THIS PAG	(When Date Entered)

REPORT DOCUMENTATION PAGE

réad instructións Before completing form Block 20 (continued)

with a microcomputer-controlled TDPAC spectrometer with logical steering that allows a standard 4-detector spectrometer to collect simultaneously 12 channels of data.

Addession For NTIS Chiai DTIC Chi spectrum is complex and both time- and sample-dependent. It is clear that in order to attain good thermal equilibrium, the low-temperature samples should be annealed consideraby longer than the tracer decay time allows. Even though the equilibrium TDPAC spectrum cannot be readily measured in the room-temperature phase, the spectra of nonequilibrium samples provide some insight into its structure.

2. Experimental apparatus and procedure

Samples were made by sealing carefully weighed quantities of 99.99% purity copper and 99.99% purity S into quartz capsules, heating above the melting point of copper, and mixing for an hour or more before cooling. Since Cu₂S is reported to dissolve small amounts of sulfur but very little copper [2], samples were made very slightly Cu-rich. This was done to ensure that the sample was nearly stoichiometric Cu₂S with small inclusions of copper. The latter is cubic and more-or-less irrelevant in a TDPAC experiment. Approximately 0.05 cm³ of this material was placed in a thin quartz sample capsule to which a few drops of llin/HCl solution were added. After drying, the capsule was sealed in vacuum and heated until the In had diffused well into the sample before any data were taken.

The TDPAC spectrometer consists of four NaI(T1) scintillation counters arranged in fixed positions at 90 degree intervals in the horizontal plane around the sample furnace. Delayed coincidences of the two gamma rays are detected at interdetector angles of 90 and 180 degrees by a variation of standard fast-slow coincidence methods. Any counter can detect the first photon of the cascade and any of the other three can detect the second. The number of coincidences for any of the twelve possible counter pairs is counted, typically for 12 to 24 hours, as a function of time t between the two gamma emissions and stored in a microcomputer memory. Sample temperature is maintained by a small nichrome-wound water-cooled furnace.

Data are analyzed by forming the ratio

$$R'(t) = \frac{2}{3} \left\{ \left(\frac{2 C_{20} C_{31} C_{02} C_{13}}{C_{10} C_{21} C_{32} C_{03} + C_{30} C_{01} C_{12} C_{23}} \right)^{1/4} - 1 \right\}$$
(1)

$$R(t) = A_{22}G_2(t) = R'(t)/[1 + 0.5R'(t)]$$
 (2)

where $C_{ij}(t)$ is the coincidence count rate when counter i detects the first and counter j the second gamma ray. Counters are numbered consecutively so C_{01} , C_{10} , C_{12} , C_{21} , etc. are 90° and C_{22} , C_{20} , C_{13} , and C_{31} are 180° coincidences. A more complete description of the spectrometer and data analysis methods will be published elsewhere.

Each type a of Cd site contributes a term,

$$R_{\alpha}(t) = A_{22} \left(S_0 + \sum_{i=1}^{3} S_i \cos(\omega_i t) e^{-(\omega_i t \delta)^2/2} \right)$$
 (3)

where the frequencies ω_i are determined by the electric field gradient (EFG), and the geometrical factors S, depend on the orientation of the EFG principal axes. The damping parameter δ arises from the spread of frequencies in imperfect crystals. The complete R(t) is the sum of such terms, weighted by the probability of occurrence.

3. Experimental results and conclusions

In the hexagonal phase, R(t) could be fit to a slightly-damped single-site axially-symmetric quadrupole interaction. Typical R(t) spectra are shown in Fig. 1. Damping was somewhat sample-dependent and increased with decreasing temperature near the transition temperature. Damping may be due to slight departures from stoichiometry at low temperature where excess sulfur is soluble. The quadrupole frequency vs. T is given in Fig. 2.

Figure 1 also shows a typical R(t) for a sample below the monoclinic/hexagonal transition temperature. Monoclinic-phase spectra could not be well fit by either a single-site or 2-site interaction with moderate damping. A "2-site" computer fit in which one site was highly damped did describe the data reasonably well however. Approximately half the nuclei are in "good" sites having a quadrupole frequency v_0 =119(6) MHz and asymmetry, parameter n=0.37(7). The highly damped portion apparently arises from nuclei in a number of different sites or from sites where the electric field gradient changes significantly during the data accumulation time.

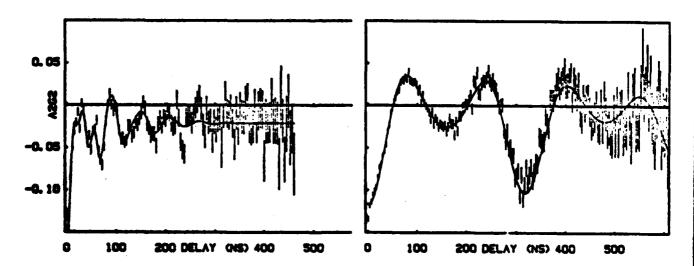


Figure 1. R(t) for dilute 111Ci impurities in Cu₂S, hexagonal phase at 136°C (right) and monoclinic phase at 100°C (left). The full lines are computer fits to (right) a single-site axially symmetric quadrupolar interaction and (left) a 2 site model with one site subject to a slightly damped nonaxially-symmetric quadrupolar interaction, and the second "site" greatly damped.

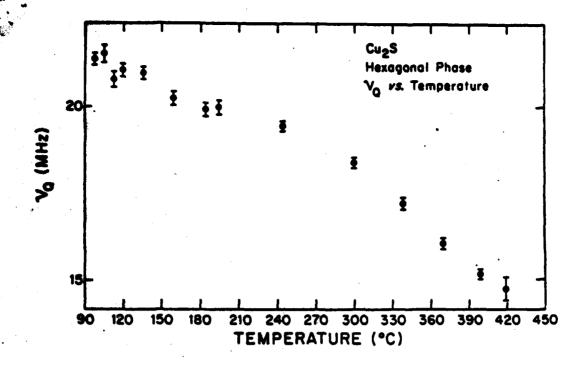


Figure 2. Quadrupolar frequency vs. T for ^{111}Cd impurities in hexagonal phase Cu_2S .

:::

ACKNOWLEDGEMENTS

This research was supported in part by the US Office of Naval Research.

REFERENCES

- [1] M. Hansen, The Constitution of Binary Alloys, (McGraw-Hill, New York, 1958) p. 620.
- [2] R. W. Potter II, Economic Geology 72 (1977) 1524.
- [3] R. W. Potter II and H. T. Evans, Jr., Jour. Research US Geol. Survey 4 (1976) 205.
- [4] H. T. Evans, Jr., Phys. Sci., 232 (1971) 69.
- [5] B. J. Mulder, Kristall und Technik 8 (1973) 825.
- (6) M. Frauenfelder and R. M. Steffen, in Perturbed Angular Correlations, ed. by E. Karlsson, E. Matthias, and K. Slegbahn (North Holland, Amsterdam, 1963).

TECHNICAL REPORT DISTRIBUTION LIST, GEN

	No.		No.
	Copies		Copies
Office of Naval Research		Naval Ocean Systems Center	
Attn: Code 413		Attn: Mr. Joe McCartney	
800 North Quincy Street		San Diego, California 92152	1
Aglington, Virginia 22217	2		
		Naval Weapons Center	
OHR Pasadena Detachment		Attn: Dr. A. B. Amster,	
Attn: Dr. R. J. Marcus		Chemistry Division	
1030 East Green Street		China Lake, California 93555	. 1
Pasadena, California 91106	1		
		Naval Civil Engineering Laboratory	
Commander, Naval Air Systems Command		Attn: Dr. R. W. Drisko	
Artn: Code 310C (H. Rosenwasser)		Port Hueneme, California 93401	1
Department of the Navy			
Washington, D.C. 20360	1	Dean William Tolles	
		Naval Postgraduate School	
Before Technical Information Center		Monterey, California 93940	1
Building 5, Cameron Station	••		
Alexandria, Virginia 22314	12	Scientific Advisor	
	• , •	Commandant of the Marine Corps	
Dr. Fred Saalfeld		(Code RD-1)	1
Chamistry Division, Code 6100		Washington, D.C. 20380	.
Havel Research Leboratory	1	Warra 1: Chila. Base anab. and Base 1 annual	
Washington, D.C. 20375	•	Naval Ship Research and Development Center	
U.S. Army Research Office		Attn: Dr. G. Bosmajian, Applied	
Actn: CRD-AA-IP		Chemistry Division	
F: 0. Son 12211		Annapolis, Maryland 21401	1
Beserri Triangle Park, H.C. 27709	1	Munchare, merarene sizai	
	•	Mr. John Boyle	•
Mr. Vincent Schaper		Materials Branch	
PENSEDC Code 2803	•	Naval Ship Engineering Center	
Ammapolis, Maryland 21402	1	Philadelphia, Pennsylvania 19112	1
		. and department, . complete control of the control	_
Meval Ocean Systems Center		Mr. A. M. Anzalone	
Arta: Dr. S. Tamamoto		Administrative Librarian	
Marine Sciences Division		PLASTEC/ARRADCOM	
Son Diego, California 91232	1	Bldg 3401	
		Dover, New Jersey 07801	1

TECHNICAL REPORT DISTRIBUTION LIST, 056

	No. Copies		No. Copies
Dr. G. A. Somorjai		Dr. W. Kohn	
Department of Chemistry		Department of Physics	
University of California		University of California	
Bertaley, California 94720	1	(San Diego)	
and and the contract of the co		La Jolla, California 92037	1
Dr. J. Marday			
Beval Research Laboratory		Dr. R. L. Park	
Surface Chemistry Division (6170)		Director, Center of Materials	
455 Overlook Avenue, S.W.		Research	
Washington, D.C. 20375	1	University of Maryland	
	_	College Park, Maryland 20742	1
Dr. J. B. Hudson		•	
Materials Division		Dr. W. T. Peria	
Rensselser Polytechnic Institute		Electrical Engineering Department	
Troy, New York 12181	1	University of Minnesota	
		Minneapolis, Minnesota 55455	1
Dr. Theodore E. Madey			
Surface Chemistry Section		Dr. Chia-wei Woo	
Department of Connerce		Department of Physics	
Mational Bureau of Standards		Northwestern University	
Hackington, D.C. 20234	1	Evanston, Illinois 60201	1
Dr. J. M. White		Dr. Robert M. Hexter	
Regertment of Chemistry		Department of Chemistry	
University of Texas		University of Minnesota	
Adetin, Texas 78712	1	Minneapolis, Minnesota 55455	1
		See B. B. Ware Statemen	
Dr. Keith R. Johnson		Dr. R. P. Van Duyne	
Department of Metallurgy and	. •	Chemistry Department	
Materials Science		Northwestern University	•
Massachusetts Institute of Technology	•	Evanston, Illinois 60201	1
Combridge, Massachusetts 02139	1	De G Gibenen	
	,	Dr. S. Sibener	
Př. J. E. Deauth		Department of Chemistry	
IM Corporation		James Franck Institute	
Thomas J. Watson Research Center		5640 Ellis Avenue	•
P. O. Box 218		Chicago, Illinois 60637	1
Torktown Heights, New York 10598	1	Do. W. A. Barrell	
		Dr. M. G. Lagally	
Dr. C. F. Flyan		Department of Metallurgical	•
Department of Physics		and Mining Engineering	
Enforcements of Illinois		University of Wisconsin	
Debane, Illinois 61801	1	Madison, Wisconsin 53706	1
A Maria Caraca C		•	

TECHNICAL REPORT DISTRIBUTION LIST, 056

Sop	o. ies	No. Copies
in Tobert Comer Printerweak of Chamistry Indea Franck Institute Spinio Rikio Avenue Skingo, Illinois 60637	Dr. K. G. Spears Chemistry Department Northwestern University Evanston, Illinois 60201	1
Thr. R. G. Wallis Supportment of Physics University of California, Irvina Electronia, California 92664	Dr. R. W. Plummer University of Pennsylvania Department of Physics Philadelphia, Pennsylvania 19104	1
Dr. V. Busisher Charlety's Department Charles Bushington Delversity Bushington; B.C. 20052	Dr. E. Yeager Department of Chemistry Case Western Reserve University Cleveland, Ohio 41106	. 1
St. 7: Barrier Threston Department Outration of Callinguisa, Salah Markan	Professor D. Hercules University of Pittsburgh Chemistry Department Pittsburgh, Pennsylvania 15260	1
Super Luchus Colletonia 19106 Pir L. C. Eller E	Professor W. Wisograd The Femorylvania State University Department of Chemistry University Park, Pennsylvania 16802	1
	Professor T. P. George The University of Rochester Glassiater Department Enchanter, New York 14627	
	Professor Dudley R. Herschbach Harvard College Office for Research Contracts 1350 Massachusetts Avenue Cambridge, Massachusetts 02138	1
Control Server Control Live Server Server Server Losses Live Server Se	Professor Horia Metiu University of California, Santa Berbera Chemistry Department So ta Barber, California 93106	
	Prof: /or a. Steckl Rensselser Polytechnic Institute Department of Electrical and Systems Engineering Integrated Circuits Laboratories 1 Troy, New York 12181	

TECHNICAL REPORT DISTRIBUTION LIST, 056

No. Copies No. Copies

Dr. John T. Yates
Department of Chemistry
University of Pittsburgh
Pittsburgh, Pennsylvania 15260

Professor G. H. Morrison Department of Chemistry Cornell University Ithaca, New York 14853

Captain Lee Myers AFOSR/NC Bolling AFB Washington, D.C. 20332

Dr. David Squire
Army Research Office
P. O. Box 12211
Research Triangle Park, NC 27709

Professor Ronald Hoffman Department of Chemistry Cornell University Ithaca, New York 14853

END

FILMED

1-84

DTIC